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Extended Calix[4]arene-Based Receptors for Molecular Recognition and Sensing

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Abstract: Recent advances in the area of recognition and sensing have shown that artificial receptors derived from extended calix[4]arenes bearing multiple $\pi$-conjugated fluorophoric or chromophoric systems have found useful to enhance binding affinity, selectivity and sensitivity for recognition and sensing of a targeted ion or molecule. A comprehensive review of various $\pi$-conjugation-extended calix[4]arene-based receptors with the highlight on the design and binding characterization for recognition and sensing is presented.

Keywords: Extended calix[4]arene; cooperative binding; recognition; sensing.

1. Introduction

There is considerable interest in developing efficient artificial receptors for molecular recognition and sensing as they play important functional roles in biological, medical, environmental, and chemical sciences [1-8]. There has been significant progress in exploring artificial receptors for sensing, particularly for metal ion sensing over past two decades [9-17]; however, design and synthesis of artificial receptors that exhibit high binding affinity, selectivity, and sensitivity to a targeted molecule or anion still pose a great challenge to the scientific community. Design of complementary structural units that exhibit recognition properties requires building in relatively strong and directional attractive forces such as hydrogen bonding, metal-ligand and electrostatic interactions. Strong binding interaction
often provides promise for high sensitivity. As a result, functional groups such as amides [18-22], ureas [23-24], thioureas [25-32], amidoureas [33-37], crown ethers [38-39], carboxylic acids, azacrown ethers [40-42], ester [43-46], and positively charged groups [47-49] have been widely used to append onto artificial receptors for recognizing neutral molecules and ions via non-covalent interactions. Receptors/sensors possessing multiple ligating groups are often found useful to promote cooperative interactions which can result in enhanced binding affinity [50]. On the other hand, a rigid and preorganized receptor site that is complementary to the target guest is also crucial to achieve high binding affinity and good selectivity.

Calix[n]arenes are macrocyclic molecules that are generally derived from the base catalyzed condensation of formaldehyde and phenol [51-56]. Such macrocycles possess a well-defined hydrophobic upper rim and hydrophilic lower rim surrounding a hollow cavity with varied dimensions that depends on the number of the phenolic units incorporated (Figure 1) [57-60].

![Molecular structure of calix[n]arene](image)

**Figure 1.** Molecular structure of calix[n]arene

Calix[n]arenes have been explored in depth as building blocks or molecular scaffolds for artificial receptors and supramolecular chemistry over the years because of their unique three-dimensional and tunable structures, together with the ease of functionalization [61-62]. The chemistry and applications of calix[n]arenes have been extensively discussed in several books [51-52, 55] and review articles [63-69]. Much research work has also been reported recently on applications of calix[n]arenes in the area of supramolecular chemistry. Calix[n]arenes can act as vase-like building blocks in the construction of artificial receptors for selective recognition of cations, anions, and neutral molecules through principles of molecular recognition. However, large-size calix[n]arenes with n > 4 (Figure 2) possess a high degree of conformational freedom, attempts to apply such calix[n]arenes (n > 4) as hosts for a specific molecular recognition have so far been limited.

As a result, calix[4]arene is of particular interest as a molecular scaffold in the design of artificial receptors because of its stable and unique three-dimensional structure. Calix[4]arene exists in four different conformations namely cone, partial-cone, 1,2-alternate, and 1,3-alternate (Figure 3).

The calix[4]arene platform can be further expanded and deepened by incorporating planar aromatic moieties at the para-positions of the phenolic units. Recently, there have been increasingly interests in taking advantage of an extended preorganized rigid platform of calix[4]arene for construction of supramolecular systems [70]. Receptors with the expanded or extended cavity are beneficial to the recognition and encapsulation with which a larger targeted guest can be accommodated. The extended π-conjugated calix[4]arene skeleton can also act as a chromophore or fluorophore for signal display. Upon binding with a guest, the change of the spectral properties of the π-conjugated skeletons would
give rise to a sensing mechanism. Furthermore, the multiple ligating groups incorporated onto the proximate multi-$\pi$-conjugated calix[4]arene assembly can act cooperatively resulting in an enhanced binding [71-74, 83].

**Figure 2.** Different sizes of calix[n]arenes

![Diagram of different sizes of calix[n]arenes](image)

**Figure 3.** Four different conformations of calix[4]arenes.

![Diagram of four different conformations of calix[4]arenes](image)

In this review, the focus is essentially upon the design and development of expanded or extended calix[4]arene-based receptors appended with multiple $\pi$-conjugated fluorophoric or chromophoric
systems at the upper rim for molecular recognition and sensing. This review also restricts the
discussion on the two conformational isomers, namely cone, and 1,3-alternate of π-conjugation-


Calix[4]arenes have been particularly attractive for their unique binding characteristics towards
alkali metal, alkaline earth, and transition metal ions. The first report on calixarene-based sensor
appeared in 1986. Diamond et al. [73] demonstrated that 1a and 1b were excellent sensors for sodium
ion. These sensors showed a potential application in determination of sodium contents in blood.
However, these receptors limited to the binding/sensing of small size of ions such as sodium ion. To
make use of calix[4]arene as a versatile molecular scaffold for recognition, much progress has been
made toward tuning the binding capability and affinity of different metal ions by means of chemical
structural modifications over the years. One of the promising approaches is to construct a preorganized
and preoriented multi-π-conjugated calix[4]arene molecular assembly bearing multiple ligating units to
enhance selective binding.

\[
\text{1a : } R = \text{H}_2\text{C} = \text{CH}_2\text{OCH}_3
\]
\[
\text{1b : } R = \text{H}_2\text{C} = \text{CH}_2\text{OCH}_2\text{CH}_3
\]

Wong et al. [74] have reported a facile and efficient protocol for the synthesis of OPP(n)-substituted
in which hexylsulfanyl end-substituted OPP(n)-substituted calix[4]arenes exhibit enhanced binding
properties for soft metal ions such as Ag\(^{+}\) and Hg\(^{2+}\).

\[
R = \text{C}_{10}\text{H}_{21} \text{ and } R' = \text{C}_6\text{H}_{13}
\]

2a : m = 2
2b : m = 3
2c : m = 4
Upon addition of Ag⁺ salt in a CDCl₃/CD₃COCD₃ (v/v = 10:1) solution of hexylsulfanyl end-capped OPP-substituted calix[4]arene assemblies, there are dramatic shifts of the proton resonances (Δδ = 0.17-0.25 ppm) and peak broadening/splitting of the methylene protons adjacent to the sulfur atom in the ¹H-NMR spectra (Figure 4).

**Figure 4.** ¹H-NMR spectra of 2b in CDCl₃ (bottom) and 2b with an addition of CF₃COOAg (top).

The binding stoichiometry of 2a·Ag⁺ determined by the Job’s plot analysis based on the fluorescence titrations and the high-resolution MALDI-TOF MS analyses consistently support a 1:1 binding mode. The binding associations of 2a and 2b toward Ag⁺ ion are one to two orders of magnitudes stronger than those of the corresponding monomeric units confirming the importance of the cooperative binding. Figure 5 shows the more stable pseudo-C₄ᵥ and less stable near the C₂ᵥ binding mode of DFT-optimized geometries of an analogous of 2a·Ag⁺ complexes.

**Figure 5.** DFT-optimized geometries of an analog of 2·Ag⁺ complex for the (a) pseudo-C₄ᵥ binding mode, and (b) near C₂ᵥ binding mode. The drawings were performed by MOLDRAW.
Wong et al. [75] have also demonstrated the use of polyether bridge at the lower rim linked between two distal phenolic rings to suppress the mobility of the cone conformer of oligophenylene-substituted 25,27-dipropoxycalix[4]crown-4s, 3a-d in order to open up the flattened cavity induced by metal complexation.

The binding interaction of Ag\(^{+}\) ion and calix[4]crown-4s was first probed by \(^1\)H NMR titrations (Figure 6) which was further confirmed with high-resolution ESI-MS or MALDI-TOF MS measurements.

**Figure 6.** \(^1\)H-NMR spectra of calix[4]crown-4s, 3a-d and the corresponding spectra upon addition of 10 equivalents of AgCF\(_3\)COO in CDCl\(_3\) and CD\(_3\)COCD\(_3\) (10:1).
They also investigated the influence of the upper rim functionality on the binding mode and site with a series of para-substituted phenyl-calix[4]crown-4s, bearing either methyl or fluoro or trimethylsilyl substituent 4-6.

The $^1$H-NMR titration studies showed that the binding was not affected by the electronic and steric properties of aryl rings substituted at the upper rim of calix[4]crown-4s. The single-crystal X-ray structure analyses further reveal that Ag$^+$ ion is bound to the ether ligands at the lower rim of calix[4]crown-4s (Figure 7). Because of the weak binding affinity of these calix[4]crown-4s 3a-d toward Ag$^+$ ion ($K_a = 30-91 \text{ M}^{-1}$), the dis-assembling process can easily be achieved by stirring the complex with KI at ambient temperature. The complexation-induced chemical shifts disappear in the $^1$H NMR spectrum upon addition of KI indicating a reversible dis-assembly. In addition, these calix[4]crown-4s have shown excellent Na$^+$/K$^+$ ion binding selectivity ($K_a (\text{Na}^+) = 1181 \text{ M}^{-1}$ and $K_a (\text{K}^+) \text{ is negligible}$) which is attributed to the better cavity size matching of sodium ion than potassium ion.

Figure 7. ORTEP drawings of (a) 3a·Ag$^+$ and (b) 3b·Ag$^+$ (25% probability). All hydrogen atoms are omitted for clarity.

Recently, Chawla et al. [76] have developed tetrathiacalix[4]arene-based receptors 7a-d by incorporating arylazo, thiazoleazo, pyridylazo, and $\beta$-naphthylazo groups onto tetrathiacalix[4]arene which exhibit Cs$^+$ and Rb$^+$ ions selectivity over other alkali and alkaline earth metal ions. They observed bathochromic shifts in the UV-vis spectroscopy as well as significant color change in the
visible region when the receptor interacted with these ions. It was believed that the cavity size of these azo-coupled tetrathiocalix[4]arenes played an important role in binding to the appropriate size of alkali metal ions. The study shows that these macrocyclic molecules can potentially be developed into useful ionic filters and sensor materials.

Chawla et al. [77] have also showed that \(8a\) and \(8b\) can detect a tiny amount (in ppm scale) of soft metal ions including \(\text{Hg}^{2+}\) and \(\text{Pd}^{2+}\) by a visual change. An immediate color change from yellow to red upon addition of \(\text{Hg(OAc)}_2\) and yellow to light purple upon addition of \(\text{K}_2\text{PdCl}_4\) were observed. Substantial shifts in the \(v_{\text{N=N}}\) and \(v_{\text{S-C}}\) stretching frequency absorption bands in the IR spectra of complexed calixarene derivatives indicate that the coordination site includes the azo moiety as well as the sulphur atom of the thioalkyl group. They pointed out that these two chromogenic calixarenes containing multi-heteroatoms with the extended azobenzene fragments show a potential to develop into a molecular diagnostic tool using the UV-vis spectroscopy.

Chawla et al. also indicated the heterocyclic calix[4]arene \(9a\) with imidazole moieties on the upper rims is highly selective extractant for \(\text{Ag}^+\) ions. IR and \(^1\text{H}\) NMR studies indicated that silver ions interacted at the upper rim and the complexation process did not involve the hydrogen atoms attached to the nitrogens of the imidazole rings. A model study using monomer derivatives \(9b\) further suggested that the size and nature of this macrocyclic structure played an important role in the recognition process.
Kim et al. [78] have reported a dual sensing probe bis-diazophenyl substituted calix[4]crown-6 10 which shows a high Pd$^{2+}$ ion selectivity over other metal ions in acetonitrile. The spectral change in the absorption spectrum upon binding arises from the blue shift of azo groups on the upper rim which is due to the positively polarized phenolic oxygen atoms of the crown loop. The Pd$^{2+}$ ion sensing was also apparent from the selective color change of this sensor from pale green to colorless. In addition, there is a fluorescence enhancement of the pyrenyl moieties via a suppressed fluorescence resonance energy transfer (FRET) upon complexation.

By elaborating on his previous work, Kim et al. [79] reported chromogenic cone-conformed tetrakis-diazophenylcalix[4]arene 11, which showed a color change from yellow to orange, upon addition of transition metal ions such as Pd$^{2+}$, Cr$^{3+}$, Fe$^{2+}$, Co$^{2+}$ and Ni$^{2+}$ ions. On the other hand, there is no significant spectral change upon the interaction with alkali and alkaline earth metal ions. The binding stoichiometry of 1:1 between 11 and Cr$^{3+}$ was determined by the Job’s experiment of continuous variation. The binding interaction is believed to be due to the release of protons from the azophenols to the quinone-hydrazone tautomer upon metal ion complexation, followed by internal chelation of metal ion to nitrogen atoms and ortho-ester carbonyl groups.

Kim et al. [80-82] have also designed and synthesized a series of extended homooxacalix[4]arene bearing tetraamides. The C-1,2-alternate conformer of tetrahomodioxacalix[4]arene N,N-dialkyl tetraamide 12a-b exhibits selective binding towards Pd$^{2+}$ ion over alkali, alkaline earth, transition metal and ammonium ions. From the X-ray structure of 12a-Pd$^{2+}$complex, it was found that Pd$^{2+}$ ion
was encapsulated by the carbonyl oxygen atoms of two adjacent amide groups and an oxygen atom of one of aryl-alkyl ethers. They also found that the C-1,3-alternate conformer 13 and the C-1,2-alternate conformer 12c gave low extractability towards Pd$^{2+}$ ion because of an intramolecular hydrogen bonding between N-H and oxygen atoms of the carbonyl or the phenyloxy groups. Their work could provide a guideline to optimize the conformation of homooxadioxacalix[4]arene derivatives for stronger metal ion binding.

![Chemical structures](image)

12a: $R = C_2H_5$, $R' = C_6H_{13}$
12b: $R = C_5H_{11}$, $R' = C_6H_{11}$
12c: $R = C_5H_{11}$, $R' = H$

Alkaline earth metal ions such as Mg$^{2+}$ and Ca$^{2+}$ are the most important metal ions in biological systems. It is necessary to differentiate one from the other in order to achieve the clinical requirements. Over the past few years, many of calix[4]arene derivatives exhibit higher selectivity for alkali metal cations than alkaline earth metal cations and only a few calix[4]arene-based Ca$^{2+}$ have been reported. Recently, attention has been drawn in using extended calix[4]arene derivatives to develop sensors for Ca$^{2+}$ ion.

Xu et al. [83] have reported the synthesis of Schiff base-linked bis-anthracene substituted calix[4]arene 14. This calix[4]arene receptor was found to show highly selectively interaction with Ca$^{2+}$ ion in MeCN-methanol mixture over other alkali and alkaline-earth metal ions with which the most intense enhancement of fluorescence of about 260 times relative to the free sensor was induced upon Ca$^{2+}$ binding. As this receptor undergoes photoinduced electron transfer (PET) from the lone-pair of the nitrogen atom of Schiff base linkage to the anthracene moiety, the fluorescence is quenched. However, the PET is inhibited and the fluorescence is restored when the lone-pair of nitrogen atom is involved in the coordination with a metal ion. The oxidation potential of the Schiff base functional group also increases substantially upon binding.
Subsequently, Kim et al. [84] also revealed bis(indoly)calix[4]crown-6, 15 in an oxidized form exhibiting selective recognition behaviors for both Ca\(^{2+}\) and F\(^{-}\) ions in CH\(_3\)CN. In the presence of F\(^{-}\) ion, the proton chemical shifts of this calixcrow\(n\)-6 unit is shifted further by an addition of Ca\(^{2+}\) ion. The results of the titration experiments of free 15 and 15·F\(^{-}\) with Ca\(^{2+}\) ion suggest that the calix[4]crown-6 binds stronger toward Ca\(^{2+}\) ion in the presence of F\(^{-}\) ion. Based on the UV-vis spectral change as well as the color change of 15, they have proposed that F\(^{-}\) ion binds to the hydrogen atoms of the expanded oxidized bisindole units in its deprotonated state by hydrogen bonds. The binding stoichiometry of 1:2 between 15 and F\(^{-}\) was found by the Job’s plot experiment. Interestingly, 15 also shows a selective color change upon addition of HSO\(_{4}\)\(^{-}\) and H\(_2\)PO\(_4\)\(^{-}\) ions in H\(_2\)O/CH\(_3\)CN (1:1 v/v). Thus, sensor 15 has a potential to be used in three combinational NOR logic gates.

Shinkai et al. [85] have reported the synthesis and alkali metal ion binding behavior of bis-pyrene substituted calix[4]crown-4, 16. Upon binding of a metal ion to the crown ether cavity, the two non-bridged distal phenol units become flattened in order to interact with the bound metal cation which would thus change the distance between two appended pyrene moieties at the upper rim. Such a change in distance would give rise to a large change in the monomer/excimer fluorescence intensity ratio. They have found that 16 bound with Na\(^{+}\) and Li\(^{+}\) but not with K\(^{+}\). In addition, they showed that the two distal pyrene moieties could interact with trinitrobenzene as there was an upfield shift (\(\Delta\delta = 0.4\) ppm) of aromatic proton resonance of trinitrobenzene in the \(^1\)H NMR spectrum upon its addition. On the other hand, the proton peak shift is slightly restored by adding the Na\(^{+}\) salt, implying the 16-trinitrobenzene complex could be dissociated through metal-binding process.
Takahashi et al. [86] have designed and synthesized novel calix[4]crown-4, 17 and calix[4]crown-4, 18 with the thienylene analogue of p-terphenoquinone as extended chromophore, which is known to exhibit strong absorption in the visible region and form a very stable radical anion and radical cation in the electrochemical reactions, for chromogenic and electrochemical sensing. Upon binding to Na⁺ or K⁺ ion, there is a dramatic conformational change of 17 and 18 from the 1,3-alternate conformer to the cone conformer which leads to a strong through-space exciton coupling of the two diametrically arranged chromophores, resulting in a significant color change. In addition, the crown-4-bridged 17 with a smaller ionophoric cavity exhibits a 100-fold higher selectivity for Na⁺ versus K⁺ than the crown-5-bridged 18. Furthermore, 17 and 18 can show not only anodic shift in the reduction potential but also cathodic shift in the oxidation potential upon addition of Na⁺ or K⁺ ion indicating that their binding behavior can be monitored electrochemically.
Rebek et al. [87-88] showed that expanded calix[4]arene tetraurea 19 in the cone conformation reversibly dimerized in apolar solvents such as CDCl₃ or CD₂Cl₂ forming well-defined capsules, as revealed by ¹H-NMR and ESI-MS spectroscopies. This dimer exhibit a large dimerization constant (Kₐ > 10⁵ M⁻¹) and are kinetically stable in the ¹H NMR time scale. The dimerization occurs via the formation of a junction of 16 C=O···N–H intermolecular hydrogen bonds (eight strong and eight weak) at the upper rim. The expanded calix[4]arene tetraurea capsule creates an egg-shaped cavity with estimated volume of ~ 400 Å³ with which it can accommodate at least two benzene- or chloroform-sized guest molecules according to the molecular modeling. This expanded capsule also forms complexes with Na⁺ cation in the apolar solvents in which the oxygen atoms of four carbonyls and four phenolic ethers at the lower rim provide enough space for tightly coordination of Na⁺ cation. The solvent exchange rate in this cavity was found to be very fast as this expanded calix[4]arene tetraurea possesses a larger room which could allow solvent molecules to pass in and out of the capsule rapidly. Although the neutral guest encapsulation is not successful, encapsulation of the positively charged molecules i-v by the dimeric capsules (19-19) is strong which is attributed to the strong ion-dipole interactions.
2.2. Extended Calix[4]arene-Based Receptors for Anions

Many biological processes in living organisms such as chemical transportation and transformation are strongly dependent of the presence or use of anions i.e. ATP and DNA. To date, anions are also significant for many industrial processes and often found as harmful pollutants such as phosphate and nitrate, which lead to entrophication. As a result, it is of utmost importance to develop useful anion sensors for different practical applications such as for cells, serum, blood, soil, and freshwater. However, it is still a great challenge to the scientific community to design and synthesize an artificial calix[4]arene-based receptors that show high binding affinity and selectivity to a targeted anion because anions have diverse topology or geometries. In addition, selective anionic receptors are relatively difficult to design and synthesize. In recent years, anion receptors based on an extended π-conjugated calix[4]arene skeleton have been developed which rely on a wide range of non-covalent interactions such as hydrogen-bonding, ion pairing or metal coordination for recognition.

Recently, Wong et al. [89] reported the synthesis and sensing properties of 20, 21, and 22 with various anions (Bu₄N⁺X⁻) including F⁻, CH₃COO⁻, H₂PO₄⁻, and PhCOO⁻ in which they were investigated by fluorescence titrations, ESI-MS measurements, and ¹H NMR spectroscopies. They have demonstrated that carbamoylphenyl-substituted calix[4]arene 20 shows greatly enhanced binding affinity and selectivity toward these anions as compared to the monomeric receptors 22, particularly for carboxylates. Amidophenyl-substituted calix[4]arene 21 also exhibits superior binding selectivity for acetate ion over fluoride ion even though it has a relatively weaker binding affinity compared to that of 20. In addition, both of these aryl-substituted calix[4]arene fluorescent receptors show much higher selectivity towards acetate over benzoate anion (for 20: Kacetate/benzoate ~ 12 and for 21: Kacetate/benzoate > 10⁴). It has been pointed out that the binding and selectivity enhancement of these arylcalix[4]arene-based sensors are attributed to the cooperative binding of the multiple ligating functional units such as carbamoyl functionalities in 20 as revealed from ab initio B3LYP 6-31G DFT quantum mechanical calculations (Figure 8). The cooperative (or bidentate) binding mode in which the hydrogen bonds formed between carboxylate and the two opposite NH groups of carbamoyl moieties flanked from the extended calix[4]arene skeleton was consistently found in the optimized structures for both 20-carboxylate complexes.
2.3. Extended Calix[4]arene-Based Receptors for Neutral Molecules

The first class of synthetic receptors for neutral molecules both in apolar solution and solid state was derived from calixarene[4]resorcinols. Recently, it has shown that the cone and 1,3-alternate conformers of calix[4]arene possess an intramolecular cavity which can host neutral guest molecules of complementary size. It has also established that the efficiency of the recognition process is strongly influenced by the rigidity of the host and the natures of the guests. Thus, calix[4]arene incorporated with extended or expanded π-conjugated framework has been developed to provide a much rigid skeleton to promote the binding and encapsulation of guests.

Fukazawa et al. [90-91] have reported the design and synthesis of extended monodeoxycalix[4]arene-based receptors 23 and 24 bearing two ethynylbenzoic acid moieties as the guest binding sites at the upper rim. The receptor 23 showed strong binding towards cyclic and acyclic ureas in CDCl$_3$ by the interactions of hydrogen bonds with a 1:1 binding stoichiometry as determined the $^1$H NMR titration experiment and the Job’s plot analysis. [90] They have also showed that the conformation of monodeoxycalix[4]crown-5, 24 can be modified upon binding of its crown ether with an appropriate alkali metal cation i.e. Na$^+$ ion. Such conformation change can lead to an enhanced binding of this receptor with the neutral cyclic and acyclic monoalkylated urea guests. [91] The enhancement is attributed to the disruption of intramolecular hydrogen bonds between the two appended benzoic acids and more favorable attractive hydrogen bonding interactions between the receptor and guest.

Rathore et al. [92] have designed and synthesized a 1,3-alternate calix[4]arene ether derivative 25 as an electron-deficient receptor upon a chemical or an electrochemical oxidation for sensing of an electron-rich guest such as nitric oxide (NO) which is vital to biological system. The 1,3-alternate calix[4]arene 25 is easily electrochemically oxidized exhibiting a low, reversible first oxidation potential in the CV measurement corresponding to the formation of a stable electron-poor cation radical and also readily chemically oxidized by MA$^+$ SbCl$_6^-$ (MA$^+$ = 9,10-dimethoxyocahydro-1,4:5,8-
dimethanoanthracene cation radical). Upon exposure of $25^+$ to gaseous NO, intense color change from bright green to dark blue is observed indicating the interaction of the two species. The $25^+$ cation radical binds to a molecule of NO with a high affinity ($K_{NO} > 10^8$ M$^{-1}$) as estimated by Venus fly trap according to the competition method. The X-ray crystallography analysis confirms that the NO molecule is deeply encapsulated inside 1,3-alternate calix[4]arene cavity.

Mendoza et al. [93] synthesized an extended calix[4]arene $26$ having the wide, coplanar $1H$-phenanthro[9,10-d]imidazol-2-yl aromatic surface situated at the upper rim. They demonstrated that an expanded cavity of $26$ could be created by hydrogen-bonded bridges or ion pairing. Upon addition of the hydrogen bond donors such as water, DMSO or hydrogen bond acceptors such as oxoanions (e.g. TFA), the slow tautomeric equilibrium, which was confirmed by splitting of the protons on the imidazole rings, as well as the well-resolved structure in solution indicated the presence of the four hydrogen-bonded or ion-paired bridges preventing the rotation of the phenanthroimidazole groups on the upper rim of the calix[4]arene platform. In the presence of monomeric $N$-methyl derivative $27$, there is a significant upfield shift of the $^1H$-NMR signal for the aromatic CH protons of $26$ in CDCl$_3$/TFA. In addition, the $N$-methyl signal of the monomer exhibits a dramatic upfield shift at 233 K, consistently, indicating that $25$ is likely inside the expanded cavity.
Subsequently, they have designed and synthesized the analogous 28 by replacing the phenanthrene moiety with 3,8-phenanthroline unit in order to introduce the heteroatoms at the corners of a square with the 90° angles for metal ion coordination. The reaction of 28 and [Re(CO)$_5$Br] gives rise to a metallocavitand 28-Re$_4$ complex [94] which has been characterized by ESI-MS, IR and NMR spectroscopies. This metallocavitand exhibits a deep, square pyramidal cavity which can fit for molecules of complementary size and shape. The binding behavior of the metallocavitand with unsubstituted calix[4]arenes and cavitands has also been investigated. It has been found that the driving force to enclose the guest inside the deep metallocavitand originates from the face-to-face stacking interactions between the electron-rich guest and the electron-deficient metallohost which was in agreement with the results of the molecular modeling and the NMR studies.
3. Conclusions

In view of the recent work, design of extended calix[4]arenes with multiple \( \pi \)-conjugated fluorophoric or chromophoric systems provides a useful strategy to enhance binding affinity, selectivity and sensitivity for recognition and sensing of a targeted ion or molecule. This review has shown that the rigid and pre-organized \( \pi \)-conjugated calix[4]arenes with multiple ligating groups and an expanded cavity can promote selectively cooperative binding for targeted guests and facilitate the encapsulation of bigger guest molecules. For practical applications, progress still needs to be made to develop water-soluble expanded calixarenes for various sensing purposes. In fact, partially water-soluble expanded calixarenes could be enough for some analytical and environmental tests. Nevertheless, expanded calixarenes still remains a largely unexplored approach to design effective and useful artificial receptors particularly for anions and biomolecules as well as supramolecular systems.

4. References


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